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2 Lessons in Chapter 15: Holt McDougal Modern Chemistry Chapter 15: Acid-Base Titration and pH. 1. The pH Scale: Calculating the pH of a Solution. Learn the history of the pH scale, how to ...

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CHAPTER 15 REVIEW Acid-Base Titration and pH SECTION 2  
SHORT ANSWER Answer the following questions in the space provided. 1. Below is a pH curve from an acid-base titration. On it are labeled three points: X, Y, and Z. a. Which point represents the equivalence point? b. At which point is there excess acid in the system? c.

## 15 Acid-Base Titration and pH

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Modern Chemistry Chapter 15 Acid-Base Titration & pH Section 1  
self-ionization of water occurs when two water molecules produce a hydronium ( $\text{H}_3\text{O}^+$ ) and a hydroxide ( $\text{OH}^-$ ) ion  
 $2 \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$   
The ionization constant ( $K_w$ ) of water is:  $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ M}$   
Acidic, Basic, & Neutral IF  $[\text{H}_3\text{O}^+] > [\text{OH}^-]$

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> [OH-] then solution is acidic.

### **Modern Chemistry Chapter 15 Acid-Base Titration & pH**

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### **Green, Michael / Modern Chemistry Textbook**

Modern Chemistry Interactive Reader ChapTER 1 Matter and Change 2 1 2 3 Chemistry Is a Physical Science 3 ... Chapter Review 469 ChapTER 15 Acid-Base Titration and pH 472 1 2 ... 2 CHAPTER 1. SECTION 1.1 Physical science is the study of nonliving objects and materials. Physics, astronomy, and geology are examples of

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Textbooks > Chemistry > Modern Chemistry: Student Edition 2012 1 > Chapter 15 > Problem 15.1.25 A nitric acid solution is found to have a pH of 2.70. Problem 15.1.25 Chapter 15

### **A nitric acid solution is found to have a pH of 2.70 ...**

Chemistry Principles of Modern Chemistry When ozone (  $O_3$  ) is placed in contact with dry, powdered KOH at  $-15^\circ C$ , the red-brown solid potassium ozonide (  $KO_3$  ) forms, according to the following balanced equation:  $5 O_3 ( g ) + 2 KOH ( s ) \rightarrow 2 KO_3 ( s ) + 5 O_2 ( g ) + H_2 O ( g )$  Calculate the volume of ozone

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needed (at a pressure of 0.134 atm and  $-15^{\circ}\text{C}$ ) to produce 4.69 g  $\text{K}_2\text{CO}_3$ .

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