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## **Chapter 13 States Of Matter**

Chapter 13 States of  
Matter137 SECTION  
13.1 THE NATURE OF  
GASES (pages  
385–389) This section  
introduces the kinetic  
theory and describes  
how it applies to gases.  
It defines gas pressure  
and explains how

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temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386)  
1.

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**STATES OF MATTER**  
**13**

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all matter consists of  
tiny particles that are  
constantly in motion  
What are the three  
assumptions of the  
kinetic theory as it  
applies to gases? -The  
particles in a gas are  
considered to be small,  
hard spheres with an



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insignificant volume.

-The motion of the particles in a gas are rapid, constant, and random.

## **Chapter 13: States of Matter Flashcards | Quizlet**

There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those

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three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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No attractive or repulsive forces exist between the particles.  
3 Chapter 13 States Of Matter Worksheet Title:  
Chapter 13 States of Matter 1 Chapter 13

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States of Matter 2

Kinetic Theory as  
Applied to Gases  
Fundamental

assumptions about  
gases. The particles in  
a gas are considered to  
be small, hard spheres  
with an insignificant  
volume.

## **Chapter 13 States Of Matter Worksheet**

→ Look at the text on  
page 315 for the  
answer. You are  
already familiar with

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the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks.

## **Chapter 13: States of Matter**

Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2

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Kinetic Theory as Applied to Gases  
Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

**PPT - Chapter 13  
States of Matter**

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**presentation ...**

Chapter 13 States of  
Matter. STUDY.

Flashcards. Learn.

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PLAY. Match. Gravity.

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bill\_brossman. Terms in  
this set (26) Kinetic

Energy. The energy an  
object has because of  
its motion. Kinetic

Theory. A theory that  
explains the states of  
matter, based on the

concept that all matter

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consists of tiny  
particles that are ...

## Answers

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Chapter 13 - States of  
*Page 15/26*

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Matter - 13.3 The  
Nature of Solids - 13.3  
Lesson Check - Page  
434: 24 Answer Solids  
have a definite shape  
and a definite volume  
because of the  
arrangement of the  
particles.

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Chemistry. Level. 11th  
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## **Chapter 13 States of Matter Flashcards**

The kinetic theory  
states that the tiny  
particles in all forms of

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matter are in constant motion! Section 13.1 The Nature of Gases Three basic assumptions of the kinetic theory as it applies to...

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Chapter 13 "States of Matter" Feb 109:15 AM

- OBJECTIVES:
- Describe the assumptions of the "kinetic theory" as it

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applies to gases. Feb

109:15 AM

•OBJECTIVES:

•Interpret gas pressure  
in terms of kinetic  
theory. Feb 109:15 AM

•OBJECTIVES: •Define  
the relationship  
between Kelvin  
temperature

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Chemistry (12th  
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Nature of Gases - 13.1  
Lesson Check - Page  
424 8 including work  
step by step written by  
community members  
like you. Textbook  
Authors: Wilbraham,  
ISBN-10: 0132525763,  
ISBN-13:  
978-0-13252-576-3,  
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**Chemistry (12th  
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States of Matter ...**

all matter consists of  
tiny particles that are

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in constant motion:  
normal boiling point:  
the boiling point of a  
liquid at a pressure of  
101.3 kPa: phase  
diagram: graph  
representing the  
relationships among  
the solid, liquid, and  
vapor states of a  
substance in a sealed  
container: standard  
atmosphere: pressure  
required to support  
760 mm of mercury ...

## **Quia - Chapter 13**

*Page 21/26*

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**"States of Matter"**

Chapter 13 - States of  
Matter - 13 Assessment  
- Page 443: 54 Answer

Temperature stays  
constant because the  
energy is needed to  
break the  
intermolecular forces  
and changing the  
phase.

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States of Matter ...**

Chapter 13 - States of  
Matter - 13.3 The

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Nature of Solids - 13.3  
Lesson Check - Page  
434: 18 Answer Atoms,  
ions, or molecules are  
packed tightly together  
in an orderly  
arrangement.

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